



Naming & Formula Writing

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Video Workbook with Dr. B

Learning how to name and write formulas for compounds comes down to:

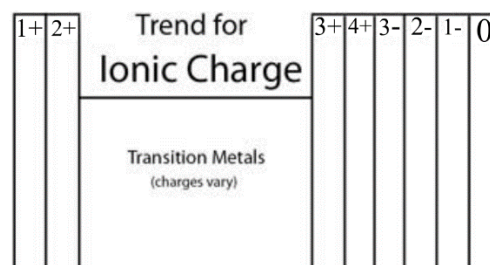
- determining the **type of compound**.
- applying the **rules** for that type of compound.
- practicing** until it becomes second nature.

Memorize these things:

Metal + Non-Metal = *Ionic*

Non-Metal + Non-Metal = *Covalent/Molecular*

Hydroxide ion: OH⁻ Nitrate: NO₃⁻ Sulfate: SO₄²⁻
 Sodium: Na Iron: Fe Silver: Ag
 Potassium: K Lead: Pb Mercury: Hg



This is the general trend for the charge ions form for elements in each Group (column).

To name or write formulas you need to know which elements are metals and non-metals.

In general metals are on the left of the Periodic Table. Non-metals are on the right.

Transition Metals are in the middle.

Hydrogen (H) is the big exception.

A key to success is practicing, making mistakes, understanding where you went wrong, and trying again.

Practice

Which of the following are non-metals?
K, Cl, H, He, Na, Al

What charge do elements in Group 2 (column 2) have?

Elements in Group 2 have a 2+ ionic charge.
 Al=3+
 K=1+, Cl=1-, H=1+, He=no charge, Na=1+,

Report errors and suggestions to DrB@breslyn.org

