



# Balancing Equations with Dr. B

Video Workbook with Dr. B.

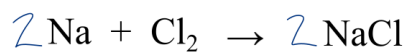
More guides at  
[www.breslyn.org](http://www.breslyn.org)

We balance equations to find the **ratios** of each substance in the equation.

When we balance equations, you can *only* change the coefficients. *Never* change the subscripts.



The goal is to have the same number of each type of atom on both sides of the equation.



$$\text{Na} = 2$$

$$\text{Cl} = 2$$

$$\text{Na} = 2$$

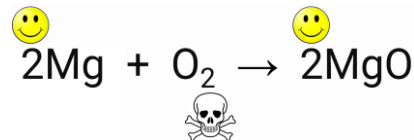
$$\text{Cl} = 2$$

The equation is balanced because there are 2 Na atoms and 2 Cl atoms in the reactants. We also have 2 Na atoms and 2 Cl atoms in the products.

Watch this video, it may be all you need to learn to balance!

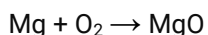
[How to Balance Equations in Five Easy Steps](#)

Remember only change the coefficients!



## Practice with Video Explanations

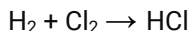
### Easy



<https://youtu.be/AZmkb9jaAml>

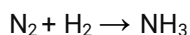


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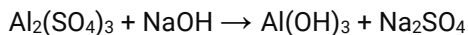
### Medium



[https://youtu.be/qf\\_aTOEa7Zs](https://youtu.be/qf_aTOEa7Zs)

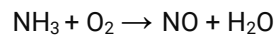


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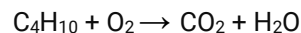
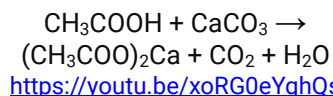


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### More Challenging



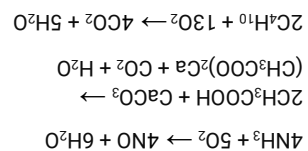
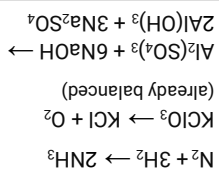
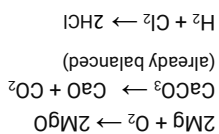
<https://youtu.be/xffqRUK1RwE>



<https://youtu.be/2oEzpBZiOm4>

Answers Below:





More practice balancing equations: [Guided Practice](#) [Practice Video](#)

## Questions people ask...

- [Why do we balance equations in the first place?](#)
- [How do I balance equations with polyatomic ions?](#)
- [How do I balance equations for combustion reactions?](#)
- [What about balancing net ionic equations?](#)

Report errors and suggestions to [DrB@breslyn.org](mailto:DrB@breslyn.org)



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